

C5 – Reactions (Exo/Endo)

	Question	Answer
1	What is an exothermic reaction?	One in which energy is given out, usually as heat or light.
2	What is an endothermic reaction?	One in which energy is taken in, usually as heat. A different example is Photosynthesis, which takes in light.
3	What happens to the temperature of the surroundings during an exothermic reaction?	It increases
4	What happens to the temperature of the surroundings during an endothermic reaction?	It decreases.
5	Give 2 examples of exothermic reactions.	Respiration, combustion, neutralisation reactions
6	Give 2 examples of endothermic reactions.	Photosynthesis & thermal decomposition
7	What does an energy level diagram show?	Energy changes during a change reaction.
8	What does Δ in an equation stand for?	Δ is the symbol for 'change' e.g. 'Change in energy' could be shown as ΔE .
9	Are exothermic reactions shown as positive or negative change in energy?	Negative
10	Are endothermic reactions shown as positive or negative change in energy?	Positive
11	What is a catalyst?	Substances that speed up chemical reactions but are unchanged at the end
12	Is bond breaking an endothermic or exothermic process?	Endothermic process.
13	Is bond making an endothermic or exothermic process?	Exothermic process
14	In terms of bonds between atoms why are exothermic reactions exothermic?	More energy is released when new bonds are made than is needed to break existing bonds
15	In terms of bonds between atoms why are endothermic reactions endothermic?	More energy is needed to break existing bonds than is released when new bonds are made